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SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope. Atomic number is the total number of protons in the nucleus of each atom of an element. 2.

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1 Matter and Change - HUBBARD'S CHEMISTRY

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SHORT ANSWER Answer the following questions in the space

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Answers: 1. Determine whether each of the following is an example of observation and data, a theory, a hypothesis, a control, or a model.
observation and data a. A research team records the rainfall in inches per day in a prescribed area of the rain forest. The

2 Measurements and Calculations

phase energy (E_{ph}) -energy stored in system due to arrangement of particles that exert attractions on one another. (energy account involved when phase changes occur)
-Attractions result in decrease of energy of a system of particles. -As particles become more tightly bound, their phase energy is lowered. -Solids possess lowest phase energy, and liquids possess more because the particles move freer.

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SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the same or different elements is the .atom 3.

3 Atoms: The Building Blocks of Matter

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the intermolecular attractions resulting

from the constant motion of electrons

and the creation of instantaneous

dipoles; forces act between all atoms

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AND molecules; forces are NOT present in intramolecular forces; increases with atomic mass or molar mass.

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