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elasticity A bottle has an opening of radius a and length b A cork of length b and radius $a + \frac{1}{2}a$ wher

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A substance combines with oxygen,
releasing a large amount of energy as heat
and light, in a(n) . 18. The decomposition
of a substance by an electric current is
called . 19. A(n) orders the elements by
the ease with which they undergo certain
chemical reactions. 20.

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Modern Chemistry 24 Chapter Test .
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continued 7. The discovery of the nucleus
was a result of Rutherford's observation
that a small percentage of the positively
charged particles bombarding the metal's
surface a. were slightly deflected as they
passed through the metal.

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and Answers 2 Measurements and

Calculations TEST B 1. a 2. c 3. c 4. a 5. c

6. c 7. a 8. a 9. time 10. mass 11. density

12. energy or work 13. length or distance

14. volume 15. area 16. qualitative 17.

quantitative 18. qualitative 19. quantitative

20. 3.00 ± 10^5 km/s 21. three 22. 2.6 ± 10

2 2 g 23. 2.50 ± 10^2 1 L 24. quantity 25.

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continued 18. When a strong acid is
titrated with a weak base, the pH of the
solution at the equivalence point is than 7.

19. A is a highly purified solid used to
check the concentration of a standard

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solution. 20. A 1 M solution of NaOH will have a pH that is

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Teacher Notes and Answers 5 The
Periodic Law TEST B 1. a 2. c 3. d 4. d 5.
a 6. a 7. c 8. a 9. lanthanides 10. 2 11.
fourth 12. transition elements 13. 32 14.
valence electrons 15. electron affinity 16.
electronegativity 17. ionization energy 18.
3 s² 3p⁴ 19. atomic radius 20 ...

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Modern Chemistry 33 Chapter Test Name
Class Date Chapter Test B, continued 15.
The energy state of an atom is called its
ground state. 16. The number of waves
that pass a point in one second is called.
17. When an electron drops from a higher-
energy state to a lower-energy state, a(n)

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Modern Chemistry 91 Chapter Test Name
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Distinguish between ionic crystals and
metallic crystals. PART V Write the
answers to the following questions on the
line to the left, and show your work in the
space provided. Modern Chemistry
Chapter 12 Review Answer Key

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Modern Chemistry 20 Chapter Test Name
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_____ 13. To calculate the number of atoms present in 2.0 mol of an element, you would a. add Avogadro's number of atoms per mole to 2.0 mole. b. subtract Avogadro's number of atoms per mole from 2.0 mole. c.

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Answers~~

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Chapter Test Chapter Test B, continued 16
Modern chemistry chapter 3 test b
answers. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

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CHAPTER 3 TEST continued Date Class
FILL IN THE BLANK Write the correct
term (or terms) in the space provided. 9, If
a particular compound is composed of
elements A and B, the ratio of the mass of
B to the mass of A will always be the
same. This is a statement of the law of
exactly 12 g of carbon-12 is referred to as
a(n) 11.

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(b) positive ions. (d) negative ions. 3. b Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is (a) higher in potential energy. (c) equal in potential energy. (b) lower in potential energy. (d) unstable. 4. b The lattice energy of compound A is greater in magnitude than that of ...

~~6 Chemical Bonding~~

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b. Explain how an atom can exist in this state. Atoms consist of a positively charged nucleus, made up of protons and neutrons, that is surrounded by a

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negatively charged electron cloud. The positive and negative charges combine to form a net neutral charge. MODERN CHEMISTRY ATOMS: THE BUILDING BLOCKS OF MATTER 19

~~3 Atoms: The Building Blocks of Matter~~

Modern Chemistry 137 Chapter Test
Name Class Date Chapter Test A,
continued _____ 19. In the figure on the previous page, the pH at the equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given. _____ 20.
In the figure on the previous page, the volume of titration standard

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CHAPTER 22 TEST Nuclear Chemistry

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Class MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice that best completes the statement or answers the question. After converting units, the nuclear mass defect is equivalent to the a. atomic mass b. electrostatic force c. energy of chemical reaction

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CHAPTER 16 REVIEW Reaction Energy

SECTION 2 SHORT ANSWER Answer

the following questions in the space

provided. 1. For the following examples, state whether the change in entropy favors the forward or reverse reaction: forward

reaction a. $\text{HCl(l)} \rightleftharpoons \text{HCl(g)}$ reverse

reaction b. $\text{C}_6\text{H}_{12}\text{O}_6\text{(aq)} \rightleftharpoons \text{C}_6\text{H}_{12}\text{O}_6\text{(s)}$ forward

reaction c. $2\text{NH}_3\text{(g)} \rightleftharpoons \text{N}_2\text{(g)} + 3\text{H}_2\text{(g)}$ forward

reaction d. $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightleftharpoons 2\text{NH}_3\text{(g)}$ forward

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