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Aufrau principle ("aufbau"means filling up). Examine the diagrams in Model 2 and the state- ments below to determine the phrase that best describes each rule. Circle the correct answer. a. Based on where a single electron is placed, the lowest potential energy electron in an atom is sublevel. found in the 3s b.

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enerw level 5 to energy level 2. Refer to Models 1 and 2 for the following questions. a. Label the picture with "n=5 to n=2" and list the corresponding color of light emitted. b. This electron transition (absorbs release energy. c. This electron moves from a (low [higher energy state to a lowe igher) energy state. d.

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the energy released in the breaking of the phosphate bonds come from when the water molecules are added to the ends of the adp and phosphate pieces, not from the actual "breaking" of phosphate bonds. $\text{adp} + \text{pi} (+ \text{energy}) \rightarrow \text{atp} (+ \text{water})$ because a phosphate group is added to the adp molecule. the reaction is endergonic. the energy released

ATP—The Free Energy Carrier

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Activities for High School Biology POGIL 15. Not all dead organisms are acted on by decomposers. Instead of being immediately recycled, the carbon from some organisms is kept in a type of long-term storage, or . carbon sink. Using Model 2, answer the questions below about this long-term storage. a. List four materials that contain this stored ...

Nutrient Cycles - cpb-us-e1.wpmucdn.com

Energy Calculations WS 2 Answer Key Assigned as HW 10/9/15- Odds only **#7 does not use $Q=mcT$ but uses $Q=mHv$ because the process of vaporization is a phase change and phase changes occur at constant temperatures as any heat being added is used towards spreading molecules apart from each other.

Piersa, Amanda / Matter and Energy

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B) Chemical Equilibrium - Pages 1&2. C) Acids, Bases, & Salts - Page 3. Worksheets. 1) Collision Theory POGIL . 2) Factors the Affect Rate . 3) Potential Energy Diagrams . 4) Dynamic Equilibrium POGIL . 5) LeChatelier's Principle POGIL . 6) LeChatelier's Principle . 7) Equilibrium Constants

AHS Chemistry Resource Site - Unit 6 - Rates & Equilibrium

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